

**Classroom expectations:**

- Be on time and prepared for class
- Show respect for yourself, others, learning, and your environment
- Bring a positive attitude and be prepared to learn
- Keep the lab clean

**Items needed for class:**

- Pen and/or pencil
- Notebook for notes and assignments – you may want to print guided notes
- Folder (optional – there will be semester finals, you will want to hold on to past chapter material!)
- Textbook
- Calculator
- Chromebook

**Late and/or incomplete assignments:**

- Handed in 1 day late – 25% grade reduction
- Handed in 2 days late – 50% grade reduction
- Handed in 3 days late – 75% grade reduction

**Discipline:**

This should not be a problem. You are here to learn and so are your classmates. If for some reason you should be removed from class, you will be given an additional assignment. Anyone caught cheating will be given a 0 on that assignment, quiz, or test.

**Cell Phones:**

You will be allowed to keep your cell phone in your bag. On test or quiz days, the phones will be placed on the counter.

**Absences:**

Students will not be punished for excused absences but are responsible for obtaining any missing notes and making up assignments and quizzes/tests. Students will have two days to make up work for every day gone, however anything handed out when the student was present is due as soon as they return. Unexcused absences will be dealt with according to the student handbook.

**Chemistry – Semester 1**

1. Understands that science is a way of knowing about the natural world and evaluating reasoning in arguments.
2. Understands that scientific inquiry uses interrelated processes to investigate the natural world.
3. Recognizes that men and women throughout history of all cultures have been involved in engineering design and scientific inquiry.
4. Understands that developments in chemistry affect society and societal concerns affect chemistry.
5. Uses physical and mathematical models to describe physical systems.
6. Obtains an in-depth understanding of how the periodic table shows how chemical and physical properties are related to atomic structure.
7. Recognizes that chemical and physical properties of matter result from the ability of atoms to form bonds.
8. Classifies and balances chemical equations and uses chemical reactions to describe chemical change.

**Chemistry – Semester 2**

1. Compares and contrasts the structure, properties, and uses of organic compounds.
2. Calculates molar mass, moles, number of particles and volume of gases.
3. Relates the properties of acids and bases to the ions they contain and predict the products of an acid-base reaction.
4. Describes and calculates the relationships in a chemical reaction, including mole/mass, mass/volume relationships, limiting reactants and percent yield.
5. Describes factors that affect the rate of a chemical reaction.

6. Understands that states of matter, properties and behavior of gases can be explained using the kinetic molecular theory